

M.H.Khan D2(H) 15-4-20,

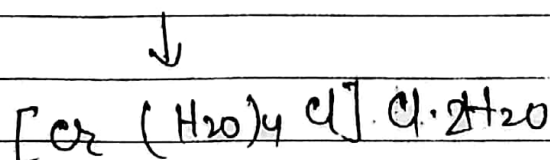
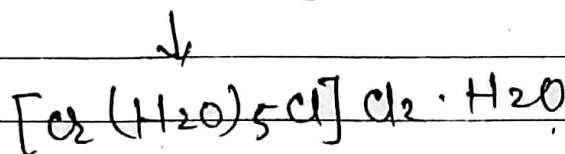
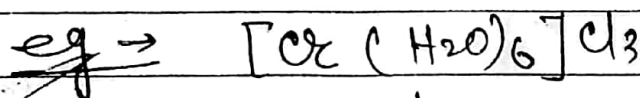
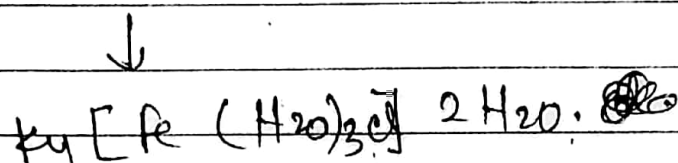
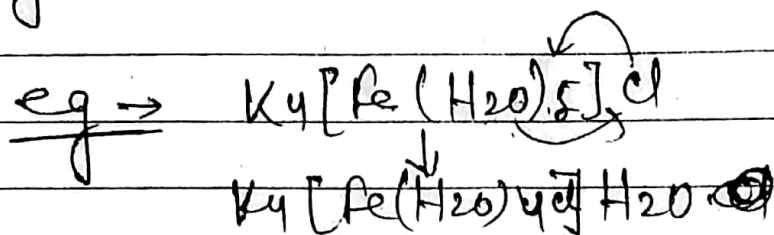
topic $\rightarrow$ Isomerism, more coordination.

classmate

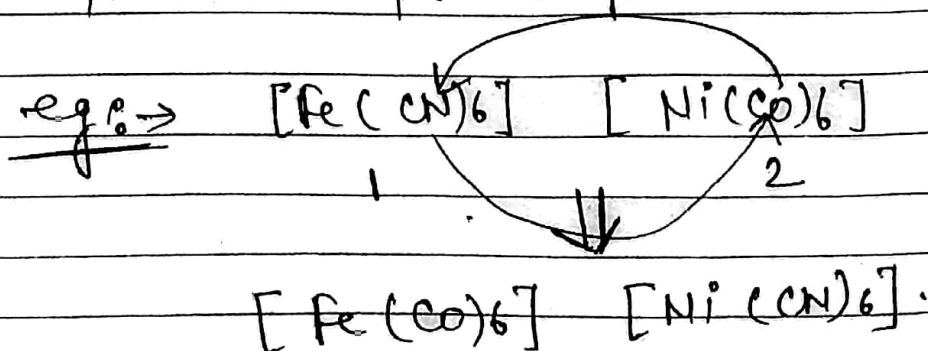
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(i) Hydrated $\rightarrow$ When compound having same Molecular formula but different str. formula due to different no. of water molecules present per [] bracket.

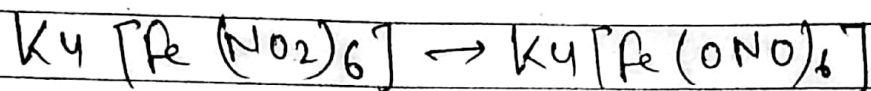


(ii) Co-ordination $\rightarrow$ When exchange of ligand occurs b/w 2 complex compound.



Topic - Isomerism, Inorganic

(IV) Linkage Isomerism: $\rightarrow$ When compound having same molecular formula but different str. formula due to different form of ligand.
eg $\rightarrow$ $\text{CN}^- \rightarrow \text{NC}^-$, $\text{NO}_2^- \rightarrow \text{ONO}^-$.
eg $\rightarrow \text{K}_4[\text{Ni}(\text{CN})_6] \rightarrow \text{K}_4[\text{Ni}(\text{NC})_6]$



$\rightarrow$ Ambident ligand show linkage isomerism.

2(a) Geometrical isomerism: $\rightarrow$ When compound having

same molecular formula but different str. formula due to different arrangement of ligand diagonally or vertically.

~~Remember~~

~~Remember~~

$\rightarrow$ Tetrahedral compound doesn't show geometrical isomerism but octahedral and square planar show geometrical isomerism.

(a) cis: $\rightarrow$ When diagonal ligand different.

Top or bottom ligand different then show geometrical isomerism.

See G.O.C. of Geometrical.